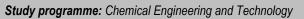


UNIVERSITY OF EAST SARAJEVO

Faculty of Technology Zvornik







583 30			Cycle I	Cycle I		Year III					
Course title			sical Chemistry 2								
Department Department			artment for Phys Inology Zvornik	sical Chem	istry, E	Electrochemical	engineering	and	materials-	Faculty	of
Course code			Cou	Course status		Semester		ECTS			
04-1-025-				Compulsory		V		7			
Teacher			ošković, PhD, full j								
Teaching assistant		•	łajić, MSc, senior								
Number of cl week)			workload (per	Individ	ual stu	dent workload semester)			Student coeffic	workload ient S。	
Lectures		ditory ercises	Laboratory exercises	Lecture	es	Auditory exercises	Laborator exercises		So		
3	0.1.1.	1	2	45		15	30	4=:4.0		33	
	3*15+	1*15+2*15	5=90 hours		400		1*15*1.33+2*	15*1.3	33)=120 hou	rs	
		A.C. C I	lotal course waning the course, st			210hours per se	emester				
Learning outcomes		 calculate heat capacities in systems based on tabular or experimentally obtained data; experimentally determine the order of the reaction and, based on the obtained solution, calculate the amounts of participants in the reaction at any moment of its development; determine coefficients of thermal conductivity, viscosity and diffusion based on experimentally obtained data and use the results in the description of industrial systems; use Faraday's laws on specific examples; experimentally use the obtained data in order to form coatings or metal powders. 									
Prerequisites		•	•				,				
Teaching meth	nods	Lectures, a	auditory and labor	atory exerc	ises, mi	d-term tests (col	lloquia).				
Syllabus ou per week	tline	2. 3. Ch 4. Irre 5. De 6. Re 7. He 8. Ca 9. Kir 10. E 11. In 12. B 13. B	Introduction to che Transfer processes temical kinetics, be eversible chemical expendence of react exections in the gas exterogeneous react talysis. Catalytic reflectrochemistry and teraction of ions in the saics of electroches in the control of the cinetics of electroches in the cinetics in	es, transfer frasic kinetic I reactions. I reactions. I reactions. It is phase. Reactions, diffus reactions, he cal reactions and electroch electrolyte emical therrodouble electrolyte electrolyte is phase.	low, vis concept Simulta on temp actions i ion conformogen s. emical de solutio modyna tric layed dation a	cosity, diffusion. s. neous reactions erature. Theory n solutions. trol of heterogen eous reactions, systems. ns. mics. er structure. nd reduction rea	of reaction spaces process heterogeneou	ses.	ctions.		
			Mid-term tests are		the 8th	week and the	15th week. Se	emeste	er verification	n is requir	ed

Obligatory reading

after the 15th week.

Author		Title, publisher	Year		Pages	
Tošković, D.		Physical Chemistry, Faculty of Technology Zvornik	1999		208-567	
		Additional reading				
Author		Title, publisher	Year			
Holclajtner-Antrunovi	ć, I.	General course of Physical Chemistry	2012	158-250		
Đorđević, S., Dražić,	V.	Physical Chemistry, Faculty of Technology and metallurgy Belgrade	2002			
Dondur, V.		Chemical kinetics, Faculty of Sciences Belgrade	2012	2 95-120		
Atkins, P.W., De Paul	a, J.	Physical Chemistry,9 th Edition, W.H. Freeman &Co., New York	2002		1-300	
Tošković, D., Aleksić,	V.	Collection of exercises in Physical Chemistry, Faculty of Technology Zvornik	2002		203-300	
Tošković, D., Vasiljev D.	ić, Lj., Lazić,	Experimental Physical Chemistry, Faculty of Technology Zvornik	2005		99-215	
Korać, F., Gutić, S. Ostojić, J., Islamović,		Practical course in physical chemistry I and II, Technological Faculty of Tuzla	2013	3 1-90		
		Type of student evaluation		Grade points	Percentag	
	Pre-exam o	bligations				
		dance	6	6%		
Obligations,		Mid-term test (colloquium) I		10	10 %	
assessment		Mid-term test (colloquium) I t Mid-term test (colloquium) II		17	17 %	
methods and			10	10 %		
grading system		theory	17	17%		
	Fig. 1 access?	rcises	10	10%		
	Final exami	/a.ral\	20	20.0/		
	Total	(oral)	30 100	30 % 100 %		
Wahnana	. 0 (0	100	100 %			
Webpage	www.tfzv.ue	25.15.Da				
Date	2023					